

## 17 Thermochemistry Answer Key

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### 17 Thermochemistry Answer Key

Answers 1.  $S(s) + O_2(g) \rightarrow SO_2(g)$   $\Delta H = -297 \text{ kJ}$  is for one mole S. We need two moles so we multiply the equation times 2:  $2 S(s) + 2 O_2(g) \rightarrow 2 SO_2(g)$   $\Delta H = -594 \text{ kJ}$  Now add in the second equation  $2 SO_2(g) + O_2(g) \rightarrow 2 SO_3(g)$   $\Delta H = -198 \text{ kJ}$  We see that the  $2 SO_2$  molecules on both the reactant side and the product side cancel, so we have  $2 S(s) + 3 O_2$

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Chapter 17 Thermochemistry 187 10. Complete the enthalpy diagram for the combustion of natural gas. Use the thermochemical equation in the first paragraph on page 517 as a guide. SECTION 17.3 HEAT IN CHANGES OF STATE (pages 520-526) This section explains heat transfers that occur during melting, freezing, boiling, and condensing.

### SECTION 17.1 THE FLOW OF ENERGY HEAT AND WORK (pages 505-510)

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162 CHAPTER 6: THERMOCHEMISTRY To convert the answer to joules, we write:  $101.3 \text{ J} \cdot 0.18 \text{ L} \cdot \text{atm} = -18 \text{ J}$  6.17 An expansion implies an increase in volume, therefore w must be  $-325 \text{ J}$  (see the defining equation for pressure-volume work.)

### Solution And Thermochemistry Test Answer Keys

$\Delta H^\circ = \Delta E^\circ + (Dn)RT$   $H_2(g) + 1/2 O_2(g) \rightarrow H_2O(l)$   $0.008314 \text{ kJ} \cdot \text{H} = 222 \text{ kJ} + (0.5) \text{ mol} \cdot 298.15 \text{ K} = 226 \text{ kJ/mol}$   $H_2$ . 3. The heat of combustion of liquid cyclohexane,  $C_6H_{12}(l)$ , is  $-3924 \text{ kJ/mole}$ . 8.25 g of cyclohexane is placed in the bomb of a bomb calorimeter with excess oxygen.

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Q. Thermochemistry is the study of changes during chemical and physical reactions. ... answer choices . True. False. Tags: Question 17 . SURVEY . 30 seconds . Q. Nonrenewable resources can be easily replaced over a short period of time. answer choices . True. False. Tags: Question 18 . SURVEY . 30 seconds . Q. Most of the energy consumed in ...

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$\Delta H^\circ = \Delta E^\circ + (Dn)RT$   $H_2(g) + 1/2 O_2(g) \rightarrow H_2O(l)$   $0.008314 \text{ kJ} \cdot \text{H} = 222 \text{ kJ} + (0.5) \text{ mol} \cdot 298.15 \text{ K} = 226 \text{ kJ/mol}$   $H_2$ . 3. The heat of combustion of liquid cyclohexane,  $C_6H_{12}(l)$ , is  $-3924 \text{ kJ/mole}$ . 8.25 g of cyclohexane is placed in the bomb of a bomb calorimeter with excess oxygen.

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