

# Read Book Chapter 9 Review Stoichiometry Section 3

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### **Chapter 9 Review Stoichiometry Section**

CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of  $N_2$  are mixed with 12.0 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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Modern Chemistry 77 Stoichiometry  
CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the answer  
on the line to the left. Show all your work  
in the space provided. 1. \_\_\_\_\_ The  
actual yield of a reaction is 22 g and the  
theoretical yield is 25 g. Calculate the  
percentage yield. 2. 6.0 mol of N<sub>2</sub> are  
mixed with 12.0 mol of H

### **CHAPTER 9 REVIEW Stoichiometry**

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Review and Chapter Summary. Learn  
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flashcards, games, and other study  
tools.

### **Chapter 9: Stoichiometry Review and Chapter Summary ...**

Stoichiometry. SECTION 1. SHORT  
ANSWER Answer the following questions  
in the space provided. 1. \_\_\_\_\_ The  
coefficients in a chemical equation  
represent the (a) masses in grams of all  
reactants and products. (b) relative  
number of moles of reactants and

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## Answer Key

products. (c) number of atoms of each element in each compound in a reaction.

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SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following

### **Modern Chemistry Stoichiometry Chapter 9 Section 1 Review ...**

CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N<sub>2</sub> are mixed with 120 mol of H

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## Answer Key Section 2 Worksheet

CHAPTER 9 REVIEW. Stoichiometry.  
SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:  
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many grams of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed? 2. Given the following equation ...

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Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  ... CHAPTER 9 REVIEW ...

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from

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the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes must be made through "File info" CorrectionKey=NL-A

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notebook, solve the following problems.  
SECTION 9.1 THE ARITHMETIC OF  
EQUATIONS Use the 3-step problem-  
solving approach you learned in Chapter  
4. 1. An apple pie needs 10 large apples,  
2 crusts (top and bottom), and 1  
tablespoon of cinnamon.

## **9 Stoichiometry Practice Problems**

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## Answer Key

CHAPTER 9 REVIEW Stoichiometry  
SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided

1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield

2 60 mol of  $N_2$  are mixed with 120 mol of  $H_2$  according to the following equation:  $N_2(g) + 3H_2(g)$

## Download Chapter 9 Review Stoichiometry Section 2 Work

Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students had an introduction to composition stoichiometry in Chapter 3 and will now move on to some more difficult problems.

## Stoichiometry Worksheet Answers Chapter 9

CHAPTER 9 REVIEW Stoichiometry

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## Answer Key

SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas:  
 $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed? ...

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