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Fundamental Environmental Chemistry

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42 Chapter 2 Environmental Chemistry Section 2.2 Stoichiometry

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43 We already know that there are 44 grams per mole of CO₂ so we do not need to recalculate that. Two moles of butane (2 mol x 58 g/mol = 116 g) yields 8 moles of carbon dioxide (8 mol x 44 g/mol = 352 g CO₂) , So we can set up the following proportion:
116 g C₄H₁₀ _ 100 g C,H,O

Environmental Chemistry 2.1 INTRODUCTION 2.2 STOICHIOMETRY ...

The field of environmental chemistry has evolved significantly since the publication of the first edition of Environmental Chemistry. Throughout the book's long life, it has chronicled emerging issues such as organochloride pesticides, detergent phosphates, stratospheric ozone depletion, the banning of chlorofluorocarbons, and greenhouse warming.

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The Handbook of Environmental Chemistry

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...

a. Describe how you would prepare 100 g of a solution that is 0.5% phenolphthalein by mass. Answer: Since the solute (phenolphthalein) is a solid, the solution is percent by mass. Mass percent means the number of grams of solute per 100 g of solution. $\text{mass percent} = (\text{mass of solute}/\text{mass of solution}) \times 100\%$.

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The text uses molecular properties such as polarity, water solubility, and vapor pressure as the starting point for understanding the environmental chemistry of various contaminants in soil, water, and the atmosphere.

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