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8. Define hydrolysis.(1 mark) 9. Write the net ionic equation for the predominant hydrolysis reaction when each of the following salts is dissolved in water. For some questions, calculations may be needed. (6 marks) a) NaF Answer \_\_\_\_\_ b) KNO<sub>2</sub> Answer \_\_\_\_\_

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HYDROLYSIS PROBLEMS 1. What is the pH of a 0.1 mol/L solution of NaCN? See the other answers in the hydrolysis problems pdf 2. Calculate the [OH<sup>-</sup>], pOH and pH of these solutions: a) 0.50 mol/L KCN b) 0.20 mol/L NaClO . Author: Gillian Ashford Created Date: 4/7/2012 1:49:56 PM ...

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Hydrolysis of  $[\text{Al}(\text{H}_2\text{O})_6]^{3+}$  Calculate the pH of a 0.10-M solution of aluminum chloride, which dissolves completely to give the hydrated aluminum ion  $[\text{Al}(\text{H}_2\text{O})_6]^{3+}$  in solution. Solution In spite of the unusual appearance of the acid, this is a typical acid ionization problem. Determine the direction of change.

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### Chemistry

Salt hydrolysis worksheet answers 1. Na<sub>2</sub>SO<sub>4</sub> Ions present: Na<sup>+</sup>, SO<sub>4</sub><sup>2-</sup> Na<sup>+</sup>: group 1 ion, therefore neutral SO<sub>4</sub><sup>2-</sup>: conjugate base of the weak acid HSO<sub>4</sub><sup>-</sup>, therefore basic SO<sub>4</sub><sup>2-</sup> + H<sub>2</sub>O HSO<sub>4</sub><sup>-</sup> + OH<sup>-</sup> Na<sub>2</sub>SO<sub>4</sub> is basic in water 2. K<sub>3</sub>PO<sub>4</sub> Ions present: K<sup>+</sup>, PO<sub>4</sub><sup>3-</sup> K<sup>+</sup>: group 1 ion, therefore neutral PO<sub>4</sub><sup>3-</sup>: conjugate base of the weak acid HPO<sub>4</sub><sup>2-</sup>, therefore basic

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Fill in only the ovals for numbers 1

through 60 on your answer sheet

Indicate all of your answers to the ...

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About This Quiz & Worksheet. When it comes to this quiz/worksheet, we're going to examine you on concepts dealing with hydrolysis reactions, what they're used for and how they're classified.

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About This Quiz & Worksheet.  
Dehydration and hydrolysis are important processes within the body. This quiz/worksheet combination will help test your understanding of the

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processes.

### **Quiz & Worksheet - Comparing Hydrolysis and Dehydration ...**

Example #1: What is the pH of a 0.0500 M solution of ammonium chloride,  $\text{NH}_4\text{Cl}$ .  $K_a$  of  $\text{NH}_4^+ = 5.65 \times 10^{-10}$ .

Solution: 1) Here is the chemical reaction (net ionic) for the hydrolysis of  $\text{NH}_4\text{Cl}$ :  $\text{NH}_4^+ + \text{H}_2\text{O} \rightleftharpoons \text{NH}_3 + \text{H}_3\text{O}^+$ . 2) Here is the  $K_a$  expression for  $\text{NH}_4^+$ :

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