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In general chemistry  
molarity is the most  
commonly used  
concentration unit: (1)  
Molarity = moles of  
solute = grams of  
solute liters of solution  
molar mass solute x

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liters of solution

Example: A student weighs 0.563 g of  $\text{FeCl}_3$  and dissolves it in enough deionized (DI) water to make 100.0 mL of solution. ( $\text{FeCl}_3$

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