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Interactive Example 16.3 - Solution

Chapter 16

Chapter 16 - Solubility and Complex Ion
Equilibria. 16.1 Solubility Equilibria and
the Solubility Product . A. Dynamic
Equilibrium $\text{CaF}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{F}^{-}(\text{aq})$
 $\text{Ca}^{2+}(\text{aq}) + 2\text{F}^{-}(\text{aq}) \rightleftharpoons \text{CaF}_2(\text{s})$

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$Ca^{2+}(aq) + 2F^{-}(aq)$ 1. Equilibrium occurs when the solution is saturated B. K_{sp} (Solubility Product Constant, Solubility Product) $K_{sp} = [Ca$

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CHAPTER 16 - Solubility and Complex Equilibria Topics to be covered:
Solubility equilibria and solubility product constant, K_{sp} ; Determination of K_{sp} from solubility and vice versa;
Factors that affect the solubility of slightly soluble salts; Solubility properties in qualitative and quantitative analyses of cations and anions;

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separation of ionic compounds based on their solubility using fractional precipitation. Complex ions and their equilibria; Effect of complex ion formation on the ...

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Chemistry 9th Edition Chapter 16 - Solubility and Complex ...

Chapter 16. Solubility and Complex Ion
Equilibria Section 16.1 Solubility
Equilibria and the Solubility Product
Solubility Equilibria Solubility product
(K_{sp}) equilibrium constant; has only one
value for a given solid at a given
temperature. Solubility an equilibrium

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position. $\text{Bi}_2\text{S}_3(\text{s}) \rightleftharpoons 2\text{Bi}^{3+}(\text{aq}) + 3\text{S}^{2-}(\text{aq})$
 $K_{\text{sp}} = [\text{Bi}^{3+}]^2 [\text{S}^{2-}]^3$

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Sections 16.1 Solubility Equilibria and the Solubility Product. 16.2 Precipitation and Qualitative Analysis ...

Ch 16 - APChemwins

Chapter 16 Solubility And Complex Ion Equilibria (2) ناردب ليعامسا د. سردم ل ا (2)
In this lecture, we continue our discussion to cover the ion product (Q),

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and determining ions concentrations in the presence of many substances. We also discuss selective precipitation.

Chapter 16 Solubility and Complex Ion Equilibria (2 ...

Section 16.3 Equilibria Involving Complex Ions
Complex Ions and Solubility
Two strategies for dissolving a

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water-insoluble ionic solid. If the anion of the solid is a good base, the solubility is greatly increased by acidifying the solution. In cases where the anion is not sufficiently basic, the

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Chapter 16: Solubility & Complex

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Ion Equilibria - ANNE ...

16.1 Solubility Equilibria and the Solubility Product Constants When a slightly soluble salt such as silver chloride, AgCl , is dissolved in water, a saturated solution is quickly obtained, because only a very small amount of the solid dissolves, while the rest remains undissolved.

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Arrhenius Concept of ... Chapter 18
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EQUILIBRIA CH 20 Ionic Equilibria III. The
Solubility Product Constant (W)

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668 . CHAPTER 16 . SOLUBILITY AND
COMPLEX ION EQUILIBRIA. Questions .
10. $MX(s) \rightleftharpoons M^{n+}(aq) + X^{n-}(aq)$ $nK_{sp} =$
 $[M^{n+}][X^{n-}]$; the K_{sp} reaction always
refers to a solid breaking u

Mrs. Beierman

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