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### Chapter 9 Stoichiometry Review Answers

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Stoichiometry b. Theoretically, how many moles of NH3 will be produced? PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N2 are mixed with 12.0 mol of H2 according to the ...

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 1. SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ The coefficients in a chemical equation represent the. (a)masses in grams of all reactants and products. (b)relative number of moles of reactants and products. (c)number of atoms of each element in each compound in a reaction.

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Chapter 9 - Stoichiometry Chapter 9 focuses on reaction stoichiometry; using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a...

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Composition stoichiometry deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry

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Stoichiometry problems can be characterized by two things: (1) the information given in the problem, and (2) the information that is to be solved for, referred to as the unknown . The given and the unknown may both be reactants, both be products, or one may be a reactant while the other is a product.