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Chemical Equilibrium
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LeChatelier's Princililie
Name: KE'i 1. In order

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to decide what effect a change in total pressure will have on an equilibrium system with gases, what is the first thing you should do when given the balanced equation?

2.,..., J. a). $N_2(g) + O_2(g) \rightleftharpoons 2NO(g)$

Answer (\D Sh', ft-

Worksheet 2-2

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Electrochemistry
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2 (Electrochemistry

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Chemistry 12 Unit 2 -

Chemical Equilibrium

Worksheet 2-2 -

LeChatelier's Principle

Page 2 6. Hydrogen

peroxide is

decomposed as

follows: $\text{H}_2\text{O}_2(\text{l}) \rightleftharpoons \text{H}_2(\text{g})$

+ $\text{O}_2(\text{g})$ $\Delta H = +187 \text{ kJ}$

Predict the direction of

equilibrium shift by

each of the following

imposed changes: a)

Increase the $[\text{H}_2]$

Answer _____

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Chemistry 12 Unit 2 -
Chemical Equilibrium
Worksheet 2-3 -

Calculations Involving
the Equilibrium

Constant Page 2 3. For
the reaction: $A_2(g) + B(g) \rightleftharpoons 2C(g)$ it is found
that by adding 1.5
moles of C to a 1.0 L
container, an
equilibrium is
established in which
0.30 moles of B are
found.

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Chemistry 12 Unit 2
- Chemical
Equilibrium

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p4 p5 p6 p7 p8 p9 p10

Answers to Unit 5

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Equilibrium, Enthalpy

and Entropy

Unit 2-Chemical

Equilibrium ...

Chemistry 12 Unit

2-Chemical Equilibrium

10. ... (Answers are

either entropy or

enthalpy) 22. As a

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reaction approaches
equilibrium, the rate of
the forward reaction ...

Unit Chemistry 12
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Reaction Kinetics
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Potential Energy
Diagrams USR TUii
POTENTIAL KNEKOY
DIAGIL4M TO ANSWER
THE QUESTIONS
BE1.0W 130-3 5 too cc,
50 o $X_2Y_2 \rightleftharpoons 2XY X_n$

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+ POOSBESS OF
REBCTION the overall
reaction as shown
exothermic or
endothermic? (2-^ do
yv\ i C 2. What is the
activation energy for
the forward reaction?
3.

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Chemistry toolkit ix
Module 5: Equilibrium
and acid reactions KEY
KNOWLEDGE 3
WORKSHEETS
WORKSHEET 5.1

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Worksheet - Unit 1 -
Foundations of
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OLI 31. (2 pta) Iridium

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has only two naturally occurring isotopes: Ir-191 with an atomic mass of 190.96058 amu, and I-193 with an atomic mass of 192.96292 amu. The atomic mass on the periodic table is 192.217 amu. Compute the percent abundance of both isotopes. 31. 31. 32.

**Solved: Page 12 Of
14 Worksheet - Unit
1 - Foundations Of**

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