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(a) = 2.25 g H₂O

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5.75×10^{-4} mol

benzoic acid 1L 122 g

benzoic acid $\times \times$ 1L

1000 mL 1 mol benzoic

acid (b) \times 37.0 mL =

2.60×10^{-3} g benzoic

acid (c) 31.7 mg SnCl₂

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$1\text{ g} \times \frac{1\text{ L}}{1000\text{ mL}} \times 4.50 = 0.143\text{ g SnCl}_2$
 $0.0225\text{ mol KBrO}_3 \times 167\text{ g KBrO}_3 \times \frac{1\text{ L}}{1000\text{ mL}} \times 11.7 = 0.270\text{ g KBrO}_3$
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