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originally organize his periodic table? By increasing atomic mass 2) What made Mendeleev's periodic table better than prior attempts to organize the elements? Mendeleev's table was able to predict the existence and properties of known elements

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## **notebook to answer the ...**

Using the periodic table, how can you identify the number of electrons in an element? The number of electrons is always equal to the number of protons, which is the atomic number. List 3 ways that the periodic table is organized. In groups, periods, and increasing atomic number.

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eight 3. nine 4. accepted 5. Dmitri  
Mendeleev 6. atomic mass 7. elements  
8. atomic number 9. Henry Moseley 10.  
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Elements and the periodic table.  
Elements are the purest form of matter and can not be broken down into any other substance by either a physical or chemical change. There are about 114

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elements and they are organized on a modern Periodic Table of the Elements. The Elements are organized by their atomic numbers from top to bottom and left to right.

## **Elements and the periodic table. 7th Grade Science ...**

Atomic Structures Study Guide

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ANSWERS 1. The Periodic Table is mostly made of metals. 2. The noble gases are the farthest group to the right on the Periodic Table. 3. Atoms having equal numbers of positive and negative charges make them neutral. 4. The three primary sub-atomic particles that atoms are made of are protons, neutrons & electrons. 5.

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The Periodic Table Questions. 1. What is ionization energy and what are its periodic trends? Explain. 2. What is electron affinity and what are its periodic trends? 3. What is an atomic radius and what are its periodic trends? Explain. 4.



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What is electronegativity and what are its periodic trends? 5.

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Elements in groups 3-12 of the periodic table. Ex: Iron, copper, nickel, silver, and gold. They are mostly hard and shiny and good conductors of electricity and

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form colorful compounds. They are less reactive than groups 1-2.

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see the teacher before the class period of the test. Key Terms Directions: Define the following terms. 1. atomic number: 2. atomic mass: 3. symbol: 4. luster: 5. ductile: 6. malleable: 7 ...

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Scientist Contribution to the

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development of the periodic table . ...  
(left) and answer these questions. a. On  
your periodic table, label the Lanthanide  
and Actinide series with your pencil. b.  
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10 points&lots of thumbs up 4 1st person

1. What is the periodic table?
2. Who developed the first periodic table?
3. How is the modern periodic table organized?
4. What are periods?
5. What are groups or families?
6. What is an element's atomic number?
8. What is an element's atomic mass?
9. Why are electrons not included in an element's

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atomic mass? 10. What is an energy level or shell ...

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20. Where are they located on the Periodic Table? On the zig-zag line 21. List all the nonmetals on the Periodic Table. He, C, N, O, F, Ne, P, S, Cl, Ar, Se,

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Br, Kr, I, Ex, At, Rn 22. Where are they located on the Periodic Table? To the right of the zig-zag line 23. List all the Lanthanides on the Periodic Table.

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## **Section 3 Reinforcement The Periodic Table Word Search ...**

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The periodic table helps predict some properties of the elements compared to each other. Atom size decreases as you move from left to right across the table and increases as you move down a column. The energy required to remove an electron from an atom increases as you move from left to right and decreases as you move down a column.

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## **Periodic Table Study Guide - Introduction & History**

Non-metals, being on the right side of the periodic table, have more protons and thus a stronger attraction to electrons in the valence shell. It's similar to how more magnets create more attraction.

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## **Explain why the periodic trends for metal and ... - study.com**

The columns in the periodic table are called the groups and the rows are called the periods. If the metals are located in the same period in the periodic table, then the metal with that highest ...

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